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1/T (K-1) 0.00180 0.00174 0.00159 0.00150 0.00143 0.00128 ln k -10.02 -8.92 -5.98 -4.26 -2.74 0.29 Gradient = -19897 = -E a 0 R E a = 19897 x 8.31 = 165000 J mol-1 = 165 kJ mol-1 6 Temperature (C) 20 50 100 150 200 250 k (mol-1 dm3 s-1) 0.0147 0.0675 0.497 2.29 7.61 20.1 1/T (K-1) ln k Gradient = -4812 = -E a R

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for many more exercises with answers TASK 1 – Bronsted-Lowry acids & bases 1 Acid = H 2O, base = NH 3 2 Acid = HCl, base = H 2O 3 Acid = HCOOH, base = KOH Chemsheets A2 1156 (QC - Kp A) © wwwCHEMSHEETScouk 3-Mar-2018 Chemsheets A2 1156 Nitrogen reacts with hydrogen as shown: 3H 2(g) + N 2(g) ⇌ 2NH 3(g) ΔH = –76 kJ mol –1 100 moles of

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Formula Sketch Shape Bond angles Metal oxidation state Co-ordination number *CHO I 1) Draw the two geometric isomers of square planar complex

TASK 1 – Drawing complexes

reacting mass calculations 1 1) Aluminium is extracted from aluminium oxide as shown. Calculate the mass of aluminium that can be formed from 1020 g of aluminium oxide.

REACTING MASS CALCULATIONS 1 - chemsheets.co.uk | 1pdf.net

initial moles 1 1 1 change in moles -0.8 +0.8 +0.8 equilibrium moles 0.2 1.8 1.8 7) CH 4(g) + H 2 O (g) CO (g) + 3 H 2(g) initial moles 1 1 1 1 change in moles -0.2 -0.2 +0.2 +0.6

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1) Increase the temperature Number of particles with E Ea more same less 2) Remove half the molecules Number of particles with E Ea more same less Energy 4) Reduce the volume of the container Number of particles with E Ea 3) Add a catalvst Number of Energy particles with E Ea more same ess Energy 6) Add an inert gas more am less Energy

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4.1 Kinetics - A-Level Chemistry

© www.CHEMSHEETS.co.uk 17-Jul-12 Chemsheets A2 029 9 IR TASK 2 1) The IR specta of six compounds are shown. The compounds are: butanoic acid butanone but-3-en-1-ol

Chemsheets A2 029 (Spectroscopy) - Weebly

1) Propene reacts with HBr to form H. H reacts with sodium hydroxide to form l, and I reacts with warm acidified potassium dichromate (VI) to form J. The infra-red spectra of H, I and J are given below, but it does indicate which is - which.

IR TASK 1 - Weebly

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Example 1 Imagine a reaction where R reacts with S to make some products: R + S -+ products If we compare experiments 1 and 2: Initial [R] Experiment mol drn-3 1.0 2.0 Init-al [S] mo drn 9.0 1.0 '2.0 The concentration of R has been doubled, but the concentration of S remains the same. he rate has oubled.

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3) In an experiment, 1.56 g of propan-1-ol ($CH_3CH_2CH_2OH$) was completely burned in air. The heat evolved raised the temperature of 0.250 dm 3 of water from 292.1 K to 339.4 K. Use this data to calculate the enthalpy of combustion of propan-1-ol (the specific heat capacity of water is 4.18 J g-1 K-1).

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This worksheet is designed to accompany chapter 14 in Chemistry: A Molecular Approach. It is a combination of other kinetics worksheets on LibreTexts.

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A rate is a measure of how some property varies with time. Speed is a familiar rate that expresses the distance traveled by an object in a given amount of time. Wage is a rate that represents the amount of money earned by a person working for a given amount of time.

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